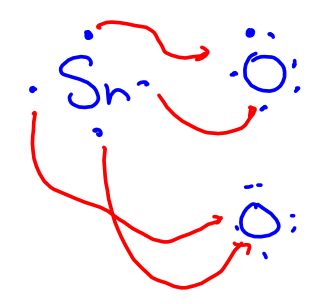
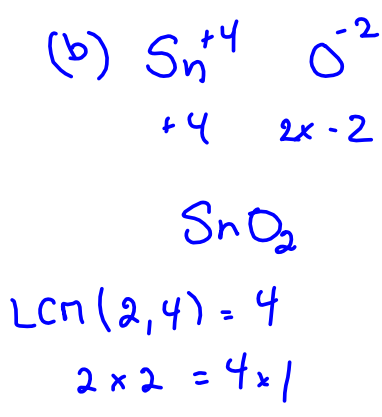
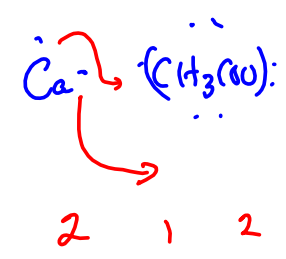
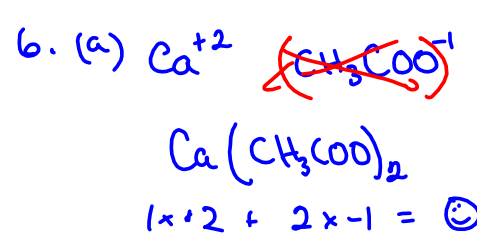
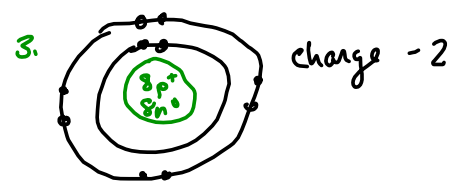
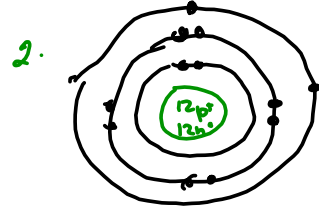
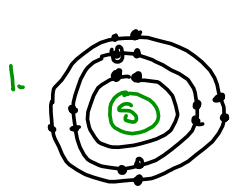


Warm up (Practice Quiz) - Using your periodic table

1. Draw the Bohr diagram for sulfur. *S, 16*
2. Draw the Bohr-Rutherford diagram for magnesium. *12, 24 #p=12 #e⁻=12 #n^o=24-12=12*
3. Draw the Bohr-Rutherford diagram for an oxygen ion. State its charge. *8, 16 8p⁺ 8e⁻ 16-8=8n^o*
4. Draw the Lewis diagram for iodine. *VII-7 I*
5. Name the following two ionic compounds: *polyatomic (back of table)*
 - a) BeF₂ *beryllium fluoride*
 - b) (NH₄)₂S *ammonium sulfide*
6. Give the chemical formulas for the following ionic compounds:
 - a) Calcium acetate
 - b) Tin (IV) oxide



Ionic Compound Naming Worksheet

Name: _____

Name the following ionic compounds or give the chemical formula for the following:

- | | |
|----------------------------------------------------------------|---------------------------------------------------------------------------------------------------------------|
| 1. <u>Na₂CO₃</u> <u>Sodium carbonate</u> | 21. sodium phosphide <u>Na⁺¹ P⁻³ Na₃P</u>
<i>need 3 +1s to balance -3</i> |
| 2. <u>NaOH</u> <u>Sodium hydroxide</u> | 22. magnesium nitrate <u>Mg⁺² NO₃⁻¹ Mg(NO₃)₂</u> |
| 3. MgBr ₂ _____ | 23. lead (II) sulfite _____ |
| 4. KCl _____ | 24. calcium phosphate _____ |
| <hr/> | |
| 5. FeCl ₂ _____ | 25. ammonium sulfate _____ |
| 6. FeCl ₃ _____ | 26. silver cyanide _____ |
| 7. Zn(OH) ₂ _____ | 27. aluminum sulfide _____ |
| 8. Be ₂ SO ₄ _____ | 28. beryllium chloride _____ |
| 9. CrF ₂ _____ | 29. copper (I) arsenide _____ |
| 10. Al ₂ S ₃ _____ | 30. iron (III) oxide _____ |
| 11. PbO _____ | 31. gallium nitride _____ |
| 12. Li ₃ PO ₄ _____ | 32. iron (II) bromide _____ |
| 13. TiI ₄ _____ | 33. vanadium (V) phosphate _____ |
| 14. Co ₃ N ₂ _____ | 34. calcium oxide _____ |
| 15. Mg ₃ P ₂ _____ | 35. magnesium acetate _____ |
| 16. Ga(NO ₂) ₃ _____ | 36. aluminum sulfate _____ |
| 17. Ag ₂ SO ₃ _____ | 37. copper (I) carbonate _____ |
| 18. NH ₄ OH _____ | 38. barium oxide _____ |
| 19. Al(CN) ₃ _____ | 39. ammonium sulfite _____ |
| 20. Be(CH ₃ COO) ₂ _____ | 40. silver bromide _____ |

Homework:

Prepare for your quiz Friday (naming ionic compounds). Work on the naming sheet.

Lab on Monday