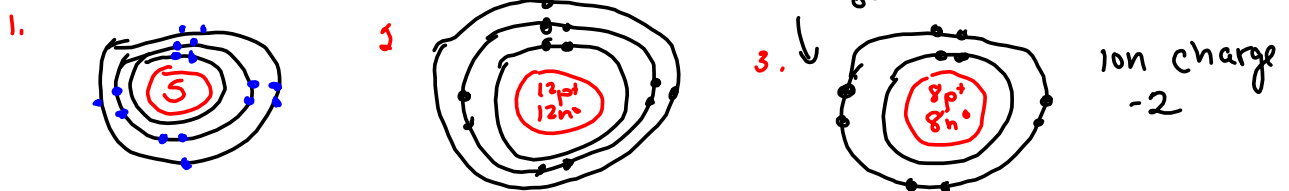


Warm up (Practice Quiz) - Using your periodic table

1. Draw the Bohr diagram for sulfur. $S, 16$
2. Draw the Bohr-Rutherford diagram for magnesium. $12, 24 \Rightarrow 12p^+, 12e^-, 24-12=12n^+$
3. Draw the Bohr-Rutherford diagram for an oxygen ion. State its charge. $8, 16 \Rightarrow 8p^+, 8e^-, 16-8=8n^+$
4. Draw the Lewis diagram for iodine.
5. Name the following two ionic compounds:
 - a) BeF_2
 - b) $(NH_4)_2S$
6. Give the chemical formulas for the following ionic compounds:
 - a) Calcium acetate
 - b) Tin (IV) oxide



4. Valence e^- symbol
 $VI\bar{I} = 7$

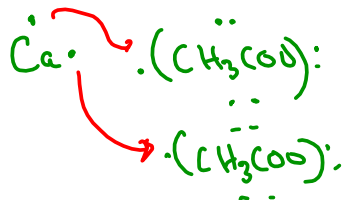
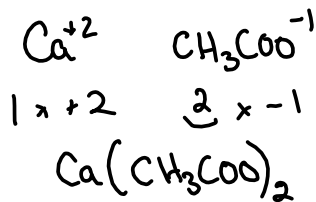


Beryllium fluoride

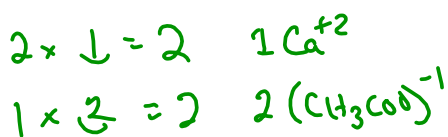


Ammonium sulfide.

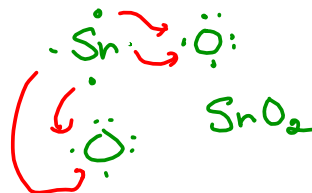
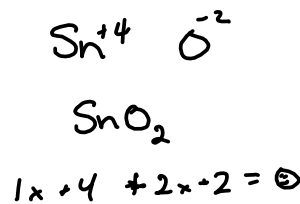
6. a) Calcium acetate



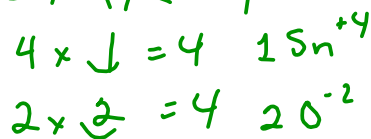
LCM 2, 1 = 2



b) Tin (IV) oxide



LCM 4, 2 = 4



Ionic Compound Naming Worksheet

Name: _____

Name the following ionic compounds or give the chemical formula for the following:

- | | |
|---|--|
| 1. $\overset{+1}{\text{Na}}_2\text{CO}_3$ <u>Sodium carbonate</u> | 21. sodium phosphide ^{phosphorus} $\overset{+1}{\text{Na}} \overset{-3}{\text{P}} \text{Na}_3\text{P}$ $3 \times +1 + -3 = 0$ |
| 2. NaOH <u>Sodium hydroxide</u> | 22. <u>magnesium nitrate</u> $\text{Mg}^{+2} \text{NO}_3^{-1} \text{Mg}(\text{NO}_3)_2$ |
| 3. MgBr_2 _____ | 23. lead (II) sulfite _____ |
| 4. KCl _____ | 24. calcium phosphate _____ |
| 5. FeCl_2 _____ | 25. ammonium sulfate _____ |
| <hr/> | |
| 6. FeCl_3 _____ | 26. silver cyanide _____ |
| 7. $\text{Zn}(\text{OH})_2$ _____ | 27. aluminum sulfide _____ |
| 8. Be_2SO_4 _____ | 28. beryllium chloride _____ |
| 9. CrF_2 _____ | 29. copper (I) arsenide _____ |
| 10. Al_2S_3 _____ | 30. iron (III) oxide _____ |
| 11. PbO _____ | 31. gallium nitride _____ |
| 12. Li_3PO_4 _____ | 32. iron (II) bromide _____ |
| 13. TiI_4 _____ | 33. vanadium (V) phosphate _____ |
| 14. Co_3N_2 _____ | 34. calcium oxide _____ |
| 15. Mg_3P_2 _____ | 35. magnesium acetate _____ |
| 16. $\text{Ga}(\text{NO}_2)_3$ _____ | 36. aluminum sulfate _____ |
| 17. Ag_2SO_3 _____ | 37. copper (I) carbonate _____ |
| 18. NH_4OH _____ | 38. barium oxide _____ |
| 19. $\text{Al}(\text{CN})_3$ _____ | 39. ammonium sulfite _____ |
| 20. $\text{Be}(\text{CH}_3\text{COO})_2$ _____ | 40. silver bromide _____ |

Homework:

Prepare for your quiz Friday (naming ionic compounds). Work on the naming sheet.

Lab on Monday